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"FUEL CELL INCORPORATING A POLYMER ELECTROLYTE MEMBRANE
GRAFTED BY IRRADIATION"

The present invention relates to a fuel cell.

5 More particularly, the present invention relates to a fuel cell incorporating a polymer electrolyte membrane grafted by irradiation, to a process for producing said polymer electrolyte membrane and to a polymer electrolyte membrane used therein.

10 The present invention moreover relates to an apparatus powered by said fuel cell.

Fuel cells are highly efficient electrochemical energy conversion devices that directly convert the chemical energy derived from renewable fuel into
15 electrical energy.

Significant research and development activities have been focused on the development of proton-exchange membrane fuel cells. Proton-exchange membrane fuel cells have a polymer electrolyte membrane disposed between a
20 positive electrode (cathode) and a negative electrode (anode). The polymer electrolyte membrane is composed of an ion-exchange polymer. Its role is to provide a means for ionic transport and for separation of the anode compartment and the cathode compartment.

25 More in particular, the traditional proton-exchange membrane fuel cells have a polymer electrolyte membrane placed between two gas diffusion electrodes, an anode and a cathode respectively, each usually containing a metal catalyst supported by an electrically conductive
30 material. The gas diffusion electrodes are exposed to the respective reactant gases, the reductant gas and the oxidant gas. An electrochemical reaction occurs at each of the two junctions (three phases boundaries) where one of the electrodes, electrolyte polymer membrane and
35 reactant gas interface.

In the case of hydrogen fuel cells, the electrochemical reactions occurring during fuel cell operation at both electrodes (anode and cathode) are the following:

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Anode: $\text{H}_2 \rightarrow 2\text{H}^+ + 2\text{e}^-$;

Cathode: $\frac{1}{2}\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{O}$;

Overall: $\text{H}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{H}_2\text{O}$.

During fuel cell operations, hydrogen permeates
5 through the anode and interact with the metal catalyst,
producing electrons and protons. The electrons are
conducted via an electrically conductive material
through an external circuit to the cathode, while the
protons are simultaneously transferred via an ionic
10 route through a polymer electrolyte membrane to the
cathode. Oxygen permeates to the catalyst sites of the
cathode where it gains electrons and reacts with proton
to form water. Consequently, the products of the proton-
exchange membrane fuel cells reactions are water,
15 electricity and heat. In the proton-exchange membrane
fuel cells, current is conducted simultaneously through
ionic and electronic route. Efficiency of said proton-
exchange membrane fuel cells is largely dependent on
their ability to minimize both ionic and electronic
20 resistivity.

Polymer electrolyte membranes play an important role
in proton-exchange membrane fuel cells. In proton-
exchange membrane fuel cells, the polymer electrolyte
membrane mainly has two functions: (1) it acts as the
25 electrolyte that provides ionic communication between
the anode and the cathode; and (2) it serves as a
separator for the two reactant gases (e.g., O_2 and H_2).
In other words, the polymer electrolyte membrane, while
being useful as a good proton transfer membrane, must
30 also have low permeability for the reactant gases to
avoid cross-over phenomena that reduce performance of
the fuel cell. This is especially important in fuel cell
applications in which the reactant gases are under
pressure and the fuel cell is operated at elevated
35 temperatures. If electrons pass through the membrane,
the fuel cell is fully or partially shorted out and the
produced power is reduced or even annulled.

Fuel cell reactants are classified as oxidants and
reductants on the basis of their electron acceptor or

electron donor characteristics. Oxidants include pure oxygen, oxygen-containing gases (e.g., air) and halogens (e.g., chlorine) and hydrogen peroxide. Reductants include hydrogen, carbon monoxide, natural gas, methane, ethane, formaldehyde, ethanol, ethyl ether, methanol, ammonia and hydrazine.

Polymer electrolyte membranes are generally based on polymer electrolytes which have negatively charged groups attached to the polymer backbone. These polymer electrolytes tend to be rather rigid and are poor proton conductors unless water is adsorbed. The proton conductivity of hydrated polymer electrolyte dramatically increases with water content.

Therefore, the proton-exchange membrane fuel cells generally require humidified gases, e.g. hydrogen and oxygen (or air), for their operations.

Among the different types of fuel cells under development, the direct methanol fuel cell (DMFC) using polymer electrolyte membranes are promising candidates for the application in portable electronic devices and in transportation (e.g. electrical vehicles).

In a direct methanol fuel cells, methanol is oxidized to carbon dioxide at the anode and oxygen is reduced at the cathode according to the following reaction scheme:

Anode: $\text{CH}_3\text{OH} + \text{H}_2\text{O} \rightarrow \text{CO}_2 + 6\text{H}^+ + 6\text{e}^-$;

Cathode: $3/2\text{O}_2 + 6\text{H}^+ + 6\text{e}^- \rightarrow 3\text{H}_2\text{O}$;

Overall: $\text{CH}_3\text{OH} + 3/2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$.

The protons are simultaneously transferred through the polymer electrolyte membrane from the anode to the cathode.

One of the major problems correlated to the use of direct methanol fuel cells is the permeation of methanol from the anode to the cathode through the membrane, a phenomenon usually known as methanol cross-over. Said methanol cross-over causes both depolarization losses at the cathode and conversion losses in terms of lost fuel. In order to improve the performance of the direct methanol fuel cell, it is necessary to eliminate or at

least to reduce said methanol cross-over. Consequently, the development of polymer electrolyte membranes which have very low permeability to methanol is desired.

5 Different types of polymer electrolyte membranes such as, for example, polyphenolsulfonic acid membranes, polystyrene sulfonate membranes, polytrifluorostyrene membranes, have been used. At present, perfluorinated membranes are the most commonly used.

10 Conventional perfluorinated membranes have a non-crosslinked perfluoroalkylene polymer main chain which contain proton-conductive functionals groups. When such membranes are ionized, the main chain is highly hydrophobic, whereas the proton-conductive side chains are highly hydrophylic. Nafion® membranes, made by
15 DuPont, are a typical example of the above mentioned membranes.

However, use of Nafion® membranes is associated with some drawbacks such as, for example, the fuel cross-over. Cross-over problems with Nafion® membranes are
20 especially troublesome in direct methanol fuel cell applications, where excessive methanol transport, which reduces efficiency and power density, occurs. Methanol cross-over not only lowers the fuel utilization efficiency but also adversely affects the oxygen cathode
25 performance, significantly lowering fuel cell performance. Moreover, the Nafion® membranes are very difficult and very expensive to be manufactured.

Various attempts have been made to provide polymer electrolyte membranes which have comparable or improved
30 properties with respect to Nafion® membranes and which are also much less expensive to be manufactured.

For example, International Patent Application WO 98/22989 discloses a polymer electrolyte membrane composed of polystyrene sulfonic acid (PSSA) and
35 poly(vinylidene fluoride) (PVDF). Said membrane may be prepared, for example, starting from the preparation of a PVDF membrane which could serve as an inert polymer matrix which is subsequently impregnated with polystyrene divinyl benzene mixtures (PS/DVB mixtures)

to produce interpenetrating polymer networks; then, the membrane so obtained is sulfonated. Instead of PVDF, other materials may be used as the inert polymer matrices such as, for example, polytetrafluoroethylene-
5 N-vinylpyrrolidone, polytetrafluoroethylene, polyvinyl-alcohol-polyacrylonitrile, polyvinyl chloride, polyvinyl alcohol, polyacrylamide, polyethylene oxide, polypropylene, polyethylene, polysulfone, sulfonated polysulfone, polyethersulfone, polyetherimide,
10 polymethylsulfoxide, polyacrylonitrile, glass membrane composites (hollow fibers), ceramic matrix host composites, zeolite matrix hosts. Said membrane is said to be particularly useful in low-temperature direct methanol fuel cell and it is said to enhance the
15 efficiency and the electrical performances of the fuel cell by decreasing methanol cross-over.

Patent Application US 2001/0026893 discloses a grafted polymer electrolyte membrane prepared by first preparing a precursor membrane comprising a polymer
20 which is capable of being graft polymerized, exposing the surface of said precursor membrane to a plasma in an oxidative atmosphere, graft-polymerizing a side chain polymer to said plasma treated precursor membrane and finally introducing a proton conductive functional group
25 to the side chain. The precursor membrane may be formed from any polymer or copolymer such as, for example, polyethylene, polypropylene, polyvinylchloride, polyvinylidenedichloride, polyvinylfluoride (PVF), polyvinilydenedifluoride (PVDF), polytetrafluoro-
30 ethylene, ethylene-tetrafluoroethylene copolymer, tetrafluoro-ethylene-perfluoroalkylvinylether copolymer, tetra-fluoroethylene-hexafluoropropylene copolymer. The side chain polymer may be any hydrocarbon polymer which contains a proton conductive functional group or which
35 may be modified to provide a proton conductive functional group. The side chain polymer may be, for example, poly(chloroalkylstyrene), poly(α -methylstyrene), poly(α -fluorostyrene), poly(p-chloromethylstyrene), polystyrene, poly(meth)acrylic acid,

poly(vinylalkylsulfonic acid), and mixtures thereof. Sulfonic acid groups are preferred as the proton conductive functional groups. The resulting grafted polymer electrolyte membrane, is said to have excellent stability and performance when used in a proton-exchange membrane fuel cell or for electrolysis of water.

Patent US 5,994,426 relates to a solid polymer electrolyte membrane which is formed of a synthetic resin which comprises (a) a main copolymer chain of a fluorocarbon-based vinyl monomer and a hydrocarbon-based vinyl monomer; and (b) a hydrocarbon-based side chain including a sulfonic group. Also disclosed is a process for producing said membrane which comprises the following steps: (a) irradiating a film-shaped copolymer made from a fluorocarbon-based vinyl monomer and a hydrocarbon-based vinyl monomer, and thereafter contacting a polymerizable alkenyl benzene with the irradiated copolymer, thereby forming a graft side chain resulting from the polymerizable alkenyl benzene; and (b) introducing a sulfonic group into the resulting graft side chain. Moreover, a modified version of said process is disclosed which comprises irradiating a film-shaped copolymer made from a fluorocarbon-based vinyl monomer and a hydrocarbon-based vinyl monomer, and thereafter contacting a polymerizable alkenyl benzene with the irradiated copolymer, thereby forming a graft side chain resulting from the polymerizable alkenyl benzene having a sulfonic group with the irradiated copolymer, thereby forming a graft side chain resulting from the polymerizable alkenyl benzene having a sulfonic group. Said membrane is said to have a high tensile strength and flexibility and it is said to be useful in polymer electrolyte fuel cell.

International Patent Application WO 00/15679 discloses a process for the preparation of a monomer-grafted cross-linked polymer comprising the steps of: (i) activating the polymer by irradiation; (ii) quenching the activated polymer so as to effect cross-linking therein; (iii) activating the cross-linked

polymer by irradiation; (iv) contacting the activated cross-linked polymer with an emulsion which comprises: (a) an unsaturated monomer; (b) an emulsifier and (c) water; for a time sufficient to effect the desired extent of grafting. Said process may be used to graft unsaturated monomers to a large number of polymers, copolymers or terpolymers formed from hydrocarbon, halogenated or perhalogenated (in particular, fluorinated or perfluorinated) monomers. Fluorinated or perfluorinated polymers, copolymers or terpolymers, are particularly preferred. Unsaturated monomers which may be used are selected from: styrene, trifluorostyrene, α -methylstyrene, α,β -dimethylstyrene, α,β,β -trimethylstyrene, ortho-methylstyrene, meta-methylstyrene, para-methylstyrene, divinylbenzene, triallylcyanurate, (meth)acrylic acid, vinylpyrrolidone, vinylpyridine, vinylacetate, trifluorovinylacetate, methyltoluene, and mixtures thereof. Said process may additionally comprises the step of sulfonating the monomer-grafted polymer. Said monomer-grafted cross-linked polymer is said to be useful in the production of non-ionic exchange membranes or ion-selective exchange membranes which can be used in various applications such as, for example, electrodialysis, dialysis, Donnan dialysis, redox cells and fuel cells.

According to the Applicant, one of the major problem encountered in fuel cells regards the performances of said fuel cells at low temperatures, e.g. at a temperatures range comprised between 20°C and 90°C. Generally, the fuel cell performances, as disclosed also in the prior art above reported, are enhanced by operating the same at higher temperatures: consequently, also the fuel cells which are said to operate at low temperatures, reach their maximum performances at high temperatures. Therefore, it would be advantageous to provide fuel cells which actually show high performances already at room temperature, e.g. at about 20°C-25°C and which retain said high performances in the whole temperatures range above reported. In addition, in the

case of direct methanol fuel cells, it is also important to minimize the methanol cross-over.

The Applicant has now found that it is possible to overcome the above mentioned problem utilizing a polymer electrolyte membrane comprising at least one polyolefin grafted by irradiation with side chains containing proton conductive functional groups, said side chains being present in a controlled amount and having a controlled length. More in particular, the Applicant has found that if the grafting irradiation process is carried out by operating at suitable conditions as reported hereinbelow, in particular at a predetermined radiation rate and for a predetermined time, it is possible to control both the amount and the length of said side chains. Said polymer electrolyte membrane is particularly useful in fuel cells operating at low temperatures, in particular at a temperatures range of from 20°C to 90°C. Said fuel cells show low cell resistance already at 20°C and retain said high performances in the whole temperatures range. Moreover, in the case of direct methanol fuel cells, said polymer electrolyte membrane shows a low methanol crossover.

According to a first aspect, the present invention thus relates to a fuel cell comprising:

- (a) an anode;
- (b) a cathode;
- (c) a polymer electrolyte membrane placed between the anode and the cathode which comprises at least one polyolefin grafted with side chains containing proton conductive functional groups;

wherein said fuel cell has:

- a value of cell resistance at 90°C not higher than 0.30 Ω cm², preferably comprised between 0.02 Ω cm² and 0.25 Ω cm², more preferably comprised between 0.05 Ω cm² and 0.20 Ω cm²;
- a value of cell resistance at 20°C differing from the value of cell resistance at 90°C of an amount not higher than 90%, preferably not higher than 70%, more preferably not higher than 50%, with respect to

the value of cell resistance at 90°C.

According to one preferred embodiment, said side chains are grafted to the polyolefin through an oxygen bridge.

5 According to one preferred embodiment, the amount of grafting [Δp (%)] of said side chains is comprised between 10% and 250%, preferably between 40% and 230%.

The amount of grafting [Δp (%)] may be calculated by the following formula:

10
$$[\Delta p (\%)] = [(W_t - W_0)/W_0] \times 100$$

wherein W_0 is the weight of the membrane before the graft polymerization reaction and W_t is the weight of the membrane after the graft polymerization reaction.

According to a preferred embodiment, said fuel cell
15 is a direct methanol fuel cell (DMFC).

For the purposes of the present description and of the claims, the expression "direct methanol fuel cell" means a fuel cell in which the methanol is directly fed into the fuel cell, without any previous chemical
20 modification, and is oxidized at the anode.

According to another preferred embodiment, said fuel cell is a hydrogen fuel cell.

According to a further aspect, the present invention relates to a polymer electrolyte membrane comprising at
25 least one polyolefin grafted with side chains containing proton conductive functional groups, said side chains being grafted to the polyolefin through an oxygen bridge.

According to one preferred embodiment, the amount of
30 grafting [Δp (%)] of said side chains is comprised between 10% and 250%, preferably between 40% and 230%.

According to a further aspect, the present invention relates to a process for producing a polymer electrolyte membrane comprising the following steps:

- 35 (i) irradiating a polyolefin in the presence of oxygen to obtain an activated polyolefin;
(ii) grafting the obtained activated polyolefin by reacting the same with at least an unsaturated hydrocarbon monomer, said hydrocarbon monomer

optionally containing at least one proton conductive functional group, to obtain side chains grafted on the activated polyolefin;

- (iii) optionally providing said grafted side chains with proton conductive functional groups, if the latter are not contained in the unsaturated hydrocarbon monomer;

wherein:

- said irradiating step (i) is carried out at a radiation rate in the range of from 0.10 Gy/s to 100 Gy/s, more preferably from 1.0 Gy/s to 10.0 Gy/s;
- said grafting step (ii) is carried out for a time period in the range of from 20 minutes to 5 hours, preferably from 30 minutes to 4 hours.

According to a further aspect, the present invention relates to an apparatus powered by the fuel cell above disclosed. Said apparatus may be an engine for vehicle transportation or, alternatively, an electronic portable device such as, for example, a mobile phone, a laptop computer, a radio, a camcorder, a remote controller.

According to one preferred embodiment, the polyolefin which may be used in the present invention may be selected from: polyethylene, polypropylene, polyvinylchloride, ethylene-propylene copolymer (EPR) or ethylene-propylene-diene terpolymer (EPDM), ethylene vinyl acetate copolymer (EVA), ethylene butylacrylate copolymer (EBA), polyvinylidenedichloride, polyvinylfluoride (PVF), polyvinylidenedifluoride (PVDF), vinylidene fluoride tetrafluoroethylene copolymer (PVDF-TFE), polyvinylidene-hexafluoropropylene copolymer, chlorotrifluoroethylene-ethylene copolymer, chlorotrifluoroethylene-propylene copolymer, polychloroethylene, ethylene-tetrafluoroethylene copolymer (ETFE), propylene-tetrafluoroethylene copolymer, propylene-hexafluoropropylene copolymer, ethylene-hexafluoropropylene copolymer. Polyethylene is particularly preferred. The polyethylene may be: high density polyethylene (HDPE) ($d = 0.940-0.970 \text{ g/cm}^3$), medium density polyethylene (MDPE) ($d = 0.926-0.940$

g/cm³), low density polyethylene (LDPE) ($d = 0.910-0.926$ g/cm³). Low density polyethylene (LDPE) is particularly preferred.

According to one preferred embodiment, the side chains may be selected from any hydrocarbon polymer chain which contains proton conductive functional groups or which may be modified to provide proton conductive functional groups. The side chains are obtained by graft polymerization of unsaturated hydrocarbon monomers, said hydrocarbon monomers being optionally halogenated. Said unsaturated hydrocarbon monomer may be selected from: styrene, chloroalkylstyrene, α -methylstyrene, α,β -dimethylstyrene, α,β,β -trimethylstyrene, ortho-methylstyrene, p-methylstyrene, meta-methylstyrene, α -fluorostyrene, trifluorostyrene, p-chloromethylstyrene, acrylic acid, methacrylic acid, vinylalkyl sulfonic acid, divinylbenzene, triallylcianurate, vinylpyridine, and copolymers thereof. Styrene and α -methylstyrene are particularly preferred.

According to a preferred embodiment, the proton conductive functional groups may be selected from sulfonic acid groups and phosphoric acid groups. Sulfonic acid groups are particularly preferred.

As already disclosed above, the present invention relates also to a process for producing a polymer electrolyte membrane.

According to one preferred embodiment, the irradiating step (i) may be carried out by γ -rays, X-rays, UV light, plasma irradiation or β -particles. γ -rays are particularly preferred.

According to one preferred embodiment, the total radiation dose in the irradiating step (i) is preferably in the range of from 0.01 MGy to 0.20 MGy, more preferably from 0.02 MGy to 0.10 MGy.

According to one preferred embodiment, after the irradiating step (i), the activated polyolefin comprises organic hydroperoxy groups ($-\text{COOH}$) in an amount of from 3×10^{-3} mol/kg to 70×10^{-3} mol/kg, preferably from 4×10^{-3} mol/kg to 50×10^{-3} mol/kg.

The amount of the organic hydroperoxy groups ($-\text{COOH}$) may be determined according to conventional techniques, e.g. by titration with a sodium thiosulfate solution.

The polyolefin may be either crosslinked or non-crosslinked before the irradiating step (i). Preferably, the polyolefin is non-crosslinked.

The activated polyolefin obtained in step (i) is stable overtime if stored at temperature of from -60°C to $+50^{\circ}\text{C}$, preferably at room temperature. Therefore, it remains activated and it is not necessary to carry out the grafting step (ii) immediately after step (i).

According to one preferred embodiment, the grafting step (ii) may be carried out at a temperature of from 15°C to 150°C , more preferably from 45°C to 55°C .

According to one preferred embodiment, the grafting step (ii) may be carried out in the presence of at least one hydroperoxy groups decomposition catalyst. Said catalyst may be selected from ferrous, cobalt, chromium or copper salts such as, for example, ferrous sulfate, ferrous ammonium sulfate, cobalt(II) chloride, chromium(III) chloride, copper chloride. Ferrous sulfate is particularly preferred. Said catalyst is preferably added in an amount of from 0.5 mg/ml to 10 mg/ml, more preferably from 1.0 mg/ml to 6.0 mg/ml.

According to one preferred embodiment, in the grafting step (ii), the hydrocarbon unsaturated monomers are dissolved in a solvent which may be selected from: ketones, such as acetone; alcohols, such as methanol; aromatic hydrocarbons, such as benzene and xylene; cyclic hydrocarbons, such as cyclohexane; ethers such as dimethylether; esters such as ethyl acetate; amides such as dimethylformamide.

According to one preferred embodiment, step (iii) may be carried out by using a sulfonating or a phosphorating agent, operating in inert-gas atmosphere, or in air. The sulfonating or phosphorating agent may be selected from: chlorosulfonic acid, fluorosulfonic acid, sulfuric acid, chlorophosphoric acid. Sulfuric acid is particularly preferred. Step (iii) may be carried out at

a temperature of from 50°C to 150°C, preferably from 70°C to 100°.

The present invention is now further illustrated with reference to the following attached figures:

5 Figure 1: is a schematic representation of a liquid feed organic fuel cell;

 Figure 2: is a graph showing cell resistance as a function of temperature;

10 Figure 3: is a schematic representation of a device used for the methanol permeation determination.

 Figure 1 shows a fuel cell (1) comprising an anode (2), a cathode (3) and the polymer electrolyte membrane (4) according to the present invention. Preferably, the anode, the cathode and the polymer electrolyte membrane
15 are integrated to form a single composite structure, with the polymer electrolyte membrane interposed between the two electrodes, commonly known as a membrane electrode assembly (MEA). Said membrane electrode assembly is usually placed in a housing which is not
20 represented in Figure 1.

 Anode (2) and cathode (3) typically comprise catalyst particles (e.g., Pt or its alloys) optionally supported on carbon particles. The catalyst particles are dispersed throughout a polymeric binder or matrix
25 which typically comprises either a proton-conductive polymer and/or a fluoropolymer. When a proton-conductive material is used, it typically comprises the same proton-conductive polymer used for the polymer electrolyte membrane. The polymeric binder or matrix
30 provides a robust structure for catalyst retention, adheres well to the polymer electrolyte membrane, aids in water management within the cell and enhances the ion exchange capability of the electrodes.

35 Anode (2) and cathode (3) are preferably formed from a platinum or from a platinum based alloy, unsupported or supported on a high surface area carbon. In the case of platinum based alloy, platinum is usually alloyed with another metal such as, for example, ruthenium, tin, iridium, osmium or rhenium. In general, the choice of

the alloy depends on the fuel to be used in the fuel cell. Platinum-ruthenium is preferable for electro-oxidation of methanol.

5 A pump (5) circulates an aqueous solution of an organic fuel in the anode compartment (6). The organic fuel is withdrawn via an appropriate outlet conduit (7) and may be recirculated. Carbon dioxide formed at the anode (2) may be vented via an outlet conduit (8) within tank (9). The fuel cell is also provided with an oxygen
10 or air compressor (10) to feed humidified oxygen or air into the cathode compartment (11).

Prior to operation, an aqueous solution of the organic fuel such as, for example, methanol, is introduced into the anode compartment (6) of the fuel
15 cell, while oxygen or air is introduced into the cathode compartment (11). Next, an external electrical load (not showed in Fig. 1) is connected between anode (2) and cathode (3). At this time, the organic fuel is oxidized at the anode and leads to the production of carbon
20 dioxide, protons and electrons. Electrons generated at the anode (2) are conducted via external electrical load to the cathode (3). The protons generated at the anode (2) migrate through the polymer electrolyte membrane (4) to cathode (3) and react with oxygen and electrons
25 (which are transported to the cathode via the external electrical load) to form water and carbon dioxide. Water and carbon dioxide produced are transported out of the cathode chamber (11) by flow of oxygen, through outlet (12).

30 Figure 3 shows a device used for the methanol permeation determination. The polymer electrolyte membrane (4) is sandwiched between a pair of graphite plates (3) provided with an array of grooves on the surface which contacts said polymer electrolyte membrane
35 (4). Said graphite plates (3) are useful in order to distribute both the methanol aqueous solution and the water evenly on the faces of the polymer electrolyte membrane (4). Said assembly [(graphite plates (3) + polymer electrolyte membrane(4))] is put between two

copper plates (2) having inlet conduits (5), (7) and outlet conduits (6), (8): the membrane is tightened by rubber gaskets. Said inlet conduits (5), (7) and outlet conduits (6), (8) flow into the graphite plates. Two tanks containing an aqueous methanol solution and distilled water respectively (not represented in Figure 3), are connected to the device (1). The aqueous methanol solution is fed [arrow (A)] through the inlet conduit (5) while water is fed [(arrow (C))] through the inlet conduit (7). One part of the aqueous methanol solution fed through the inlet conduit (5) passes through the membrane (4) while the remaining part comes out [(arrow (B))] from the outlet conduit (6). The aqueous methanol solution which passes through the membrane (4) mixed with the water fed [(arrow (C))] through the inlet conduit (7) comes out [(arrow (D))] from the outlet conduit (8).

The methanol permeation is determined by gas-chromatographic analysis of the aqueous methanol solution recovered both from the outlet conduits (6) and the outlet conduit (8), [arrow (B)] and [arrow (D)] respectively.

The present invention will be further illustrated hereinbelow by means of examples.

25 EXAMPLE 1

A low density polyethylene (LDPE) film was irradiated by γ -rays at a total radiation dose of 0.05 MGy, at a radiation rate of 5.2 Gy/s, from a ^{60}Co -irradiation source, in air, at room temperature.

30 Styrene (purity $\geq 99\%$) from Aldrich was washed with an aqueous solution of sodium hydroxide at 30% and then washed with distilled water until the wash water had a neutral pH. The so treated styrene was then dried over calcium chloride (CaCl_2) and was distilled under reduced pressure.

Then, using the styrene purified as above, a styrene/methanol solution (60:40 vol.%) containing 2 mg/ml of ferrous sulfate ($\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$) was prepared.

The irradiated LDPE film was immersed in 100 ml of

the styrene/methanol solution prepared as above using a reaction vessel equipped with a reflux condenser. The reaction vessel was then heated in a water bath until boiling of the solution.

5 After 1 hour (grafting time) the LDPE film was removed from the reaction vessel, washed with toluene and methanol three times, then dried in air and vacuum at room temperature to constant weight.

Then, the grafted LDPE film was immersed in a
10 concentrated sulfuric acid solution (96%) and heated for 2 hours at 98°C in a glass ampoule supplied with reflux condenser. Thereafter, the LDPE film was taken out of the solution, was washed with different aqueous solutions of sulfuric acid (80%, 50% and 20%
15 respectively), and finally with distilled water until the wash water had a neutral pH. Then, the film was dried in air at room temperature and after in vacuum at 50°C to constant weight obtaining a membrane according to the present invention.

20 EXAMPLE 2

A membrane was prepared as disclosed in Example 1 the only difference being the grafting time: 2 hours.

EXAMPLE 3

A membrane was prepared as disclosed in Example 1
25 the only difference being the grafting time: 4 hours.

EXAMPLE 4

The membranes obtained as disclosed in the above Examples 1-3, were subjected to the following characterizations.

30 (a) Determination of the amount of the organic hydroperoxy groups after irradiation

The determination of the amount of the organic hydroperoxy groups after irradiation, was carried out as follows.

35 2 g of the irradiated polymer were added to 10 ml of chloroform in a flask and were maintained under stirring until complete dissolution.

15 ml of acetic acid and 1 ml of potassium iodine were then added. The flask was rapidly closed,

maintained under stirring for 1 min at room temperature and, subsequently, for 5 min in the dark at a temperature comprised between 15°C and 25 °C. Then, 75 ml of distilled water were added. The released iodine was then titrated with a 0.002 N sodium tiosulfate solution, under vigorous stirring, using a starch solution as indicator. At the same time a standard was titrated.

The amount of organic hydroperoxy groups (-COOH), expressed in mol of active oxygen per kg of polymer (mol/kg), was calculated according to the following formula:

$$(-\text{COOH}) \text{ groups} = 32(V \cdot N/m) \cdot 1000$$

wherein V (expressed in ml) is the volume of the standard sodium tiosulfate solution, after correction with the standard, N is the normal concentration of the sodium tiosulfate solution and m is the weight of the analyzed polymer.

The obtained results are given in Table 1 and are the arithmetical average value of two measurements.

(b) Determination of the amount of grafting [Δp (%)]

The amount of grafting [Δp (%)] was calculated as disclosed above: the obtained results are given in Table 1.

(c) Determination of ion-exchange capacity (IEC) after sulfonation

The ion-exchange capacity (IEC) was determined as follows.

The membranes obtained as disclosed in the above Examples 1-3 were immersed in 1 N HCl aqueous solution, at room temperature, for 1 hour, in order to obtain the samples in the protonic form. Thereafter, the membranes were washed with deionised water at 50°C-60°C and were dried in oven at 80°C under vacuum for 2 hours.

The membranes were then immersed in a 1M NaCl solution for 1 h, in order to exchange the hydrogen ions with sodium ions. The hydrogen ions which passed through the membranes were titrated by neutralization with an 0.01 N NaOH aqueous solution in order to determine the

ion-exchange capacity of the membranes. The obtained results, expressed in milli-equivalent/g, are given in Table 1.

TABLE 1

EXAMPLE	(-COOH) (mol/kg)	Grafting time (h)	[Δp (%)]	IEC (milli- equivalent/g)
1	4.6×10^{-3}	1	73.5	2.05
2	4.6×10^{-3}	2	100	2.36
3	4.6×10^{-3}	4	220	3.51

EXAMPLE 5

Cell resistance measurement

Fuel cell electrodes ELAT type commercialized by E-TEK Inc. (Somerset, N.J.), were used to obtain a membrane electrode assembly (MEA). The carbon electrodes contained Pt in an amount of 0.5 mg/cm^2 both for the anode and the cathode. The electrodes were put into contact with the membrane each at opposite faces of the membrane and the MEA assembly so obtained was installed in a fuel cell housing that was tightened at 1 kg/cm^2 pressure.

The geometrical electrode area of the electrode/membrane assembly was 5 cm^2 . The MEA assembly was installed in a single cell test system which was purchase by Glob Tech Inc. The system was composed of two copper current collector end plates and two graphite plates containing rib channel patterns allowing the passage of an aqueous solution to the anode and humidified oxygen to the cathode. The single cell was connected to an AC Impedance Analyser type 4338B from Agilent. The fuel cell so constructed was operated at different temperatures in a range comprised between 20°C and 90°C . Water was supplied to the anode through a peristaltic pump and a preheater maintained at the cell

temperature. Humidified oxygen was fed to the cathode at atmospheric pressure. The oxygen humidifier was maintained at a temperature 10°C above the cell temperature. The operating conditions simulated those of direct methanol fuel cell (DMFC). Cell resistance was measured at the fixed frequency of 1 KHz and under an open circuit by means of the AC Impedance Analyser above reported operating in the temperatures range of from 20°C to 90°C.

After inserting the MEA assembly into the single test housing, the cell was equilibrated by distilled water and humidified oxygen. After obtaining a constant value of resistance, the cell was heated up to 90°C stepwise and resistance measurements, expressed in $\Omega \text{ cm}^2$, were carried out at different temperatures.

The tested membranes were the following:

- Nafion® 112 by Dupont (50 μm thickness);
- Nafion® 117 by Dupont (170 μm thickness);
- membrane obtained according to Example 3 (20 μm thickness).

The obtained results are given in Table 2 and in Fig. 2.

In Table 2 was also reported the percentage difference (R%) between the value of cell resistance at 20°C and the value of cell resistance at 90°C with respect to the value of cell resistance at 90°C according to the following formula:

$$(R\%) = [(R_{20^\circ\text{C}} - R_{90^\circ\text{C}}) / R_{90^\circ\text{C}}] \times 100$$

wherein $R_{20^\circ\text{C}}$ is the value of cell resistance at 20°C and $R_{90^\circ\text{C}}$ is the value of cell resistance at 90°C.

Table 2 and Fig. 2 clearly show that the fuel cell having the membranes according to the present invention (Example 3) has a high performance already at low temperatures (20°C) and maintain said high performances in the whole temperatures range.

TABLE 2

TEMPERATURE (°C)	CELL RESISTANCE (Ω cm ²)		
	Nafion® 112	Nafion® 117	Example 3
20	0.230	0.540	0.090
25	-	-	0.088
30	0.200	0.460	-
35	-	-	0.081
40	0.195	0.360	0.080
50	0.165	0.330	0.075
60	0.140	0.280	0.071
70	0.125	0.240	0.067
80	0.115	0.220	0.065
90	0.110	0.190	0.061
(R%)			
	109	184	47.5

Methanol permeation determination

5 The methanol permeation determination was carried out according to the method described above using a device schematically represented in Fig. 3. The membranes utilized are those of Table 3.

10 Two tanks of equal volume (200 ml) containig a 2M methanol solution and distilled water were connected to the device through two peristaltic pumps (not represented in Figure 3): the flow speed of the methanol and of the distilled water to the inlet conduits (5) and (7) respectively, was 1.92 ml/min.

Aliquots of 2.4 ml of the outlet solutions both from the outlet conduit (6) and the outlet conduit (8), [(arrow (B))] and [(arrow (D))] respectively, were taken after 15 min and 200 μ l of the same were analysed by means of a cromathograph VEGA Series 2 GC6000 from Carlo Erba equipped with a Carbopack 3% SP 1500 column and a flame ionization detector at 80°C. As a standard a 100 ppm aqueous methanol solution was used. The obtained results are given in Table 3.

TABLE 3

EXAMPLE	l (μ m)	p (mol min ⁻¹ cm ⁻¹)	p* (mol min ⁻¹ cm ⁻¹)
1	55	0.56×10^{-6}	3.08×10^{-9}
2	60	0.57×10^{-6}	3.42×10^{-9}
3	20	0.97×10^{-6}	1.94×10^{-9}
Nafion® 112	50	2.14×10^{-6}	1.07×10^{-8}
Nafion® 117	170	7.78×10^{-7}	1.32×10^{-8}

(l): membrane;

(p): methanol permation rate;

(p*): methanol permation rate normalized to the membrane thickness.

The data above reported show that the permeation rate of the membranes according to the present invention (Examples 1-3) are lower than the permation rate of the membranes of the prior art (Nafion® 112 and Nafion® 117).